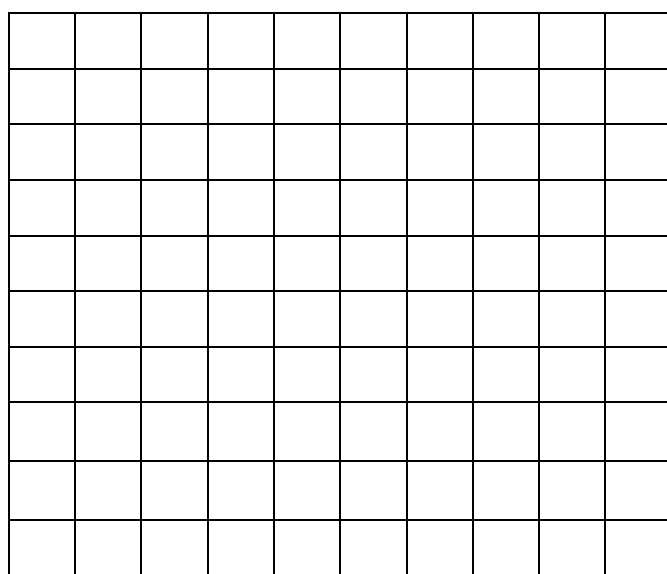




- a. In a full sentence, state the trend between periods and atomic radius.
- b. Use models of Bohr atoms from your period to represent the trend in atomic radius for Periods.

3. Using your reference table S, graph the Atomic Number versus Ionization Energy for Group \_\_\_\_.

<b>Atomic Number</b>	<b>Ionization Energy</b>



- a. In a full sentence, state the trend between groups and ionization energy.
- b. Use Bohr models for atoms in your group to represent the trend in ionization energy for groups.



- a. In a full sentence, state the trend between groups and electronegativity.
- b. Use Bohr models for atoms in your group to represent the trend in electronegativity for groups.

6. Using your reference table S, graph the Atomic Number versus Electronegativity for Period \_\_\_\_.

Atomic Number	Electronegativity


- a. In a full sentence, state the trend between periods and electronegativity.
- b. Use Bohr models for atoms in your period to represent the trend in electronegativity for periods.