**Name : \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ Periodic Table Graphing Lab**

Procedure: For all graphs, label the axes, use constant intervals, circle your points, and connect with a best fit line. Use pencil and a ruler. **Your graphs will be graded on both accuracy and appearance.**

1. Using your reference table S, graph the Atomic Number versus Radius for the Group \_\_\_\_.

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| |  |  | | --- | --- | | **Atomic Number** | **Radius** | |  |  | |  |  | |  |  | |  |  | |  |  | |  |  | |  |  | |  |  |  |  |  |  |  |  |  |
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* 1. In a full sentence, state the trend between groups and atomic radius.
  2. Use Bohr models of atoms in your group to represent the trend in atomic radius trend for groups.

1. Using your reference table S, graph the Atomic Number versus Radius for Period \_\_\_\_.

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| |  |  | | --- | --- | | **Atomic Number** | **Radius** | |  |  | |  |  | |  |  | |  |  | |  |  | |  |  | |  |  | |  |  | |  |  |  |  |  |  |  |  |  |
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* 1. In a full sentence, state the trend between periods and atomic radius.
  2. Use models of Bohr atoms from your period to represent the trend in atomic radius for Periods.

1. Using your reference table S, graph the Atomic Number versus Ionization Energy for Group \_\_\_\_.

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| |  |  | | --- | --- | | **Atomic Number** | **Ionization Energy** | |  |  | |  |  | |  |  | |  |  | |  |  | |  |  | |  |  | |  |  |  |  |  |  |  |  |  |
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* 1. In a full sentence, state the trend between groups and ionization energy.
  2. Use Bohr models for atoms in your group to represent the trend in ionization energy for groups.

1. Using your reference table S, graph the Atomic Number versus Ionization Energy for Period \_\_\_\_.

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| |  |  | | --- | --- | | **Atomic Number** | **Ionization Energy** | |  |  | |  |  | |  |  | |  |  | |  |  | |  |  | |  |  | |  |  | |  |  |  |  |  |  |  |  |  |
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* 1. In a full sentence, state the trend between periods and ionization energy.
  2. Use Bohr models for atoms in your period to represent the trend in ionization energy for periods

1. Using your reference table S, graph the Atomic Number versus Electronegativity for Group \_\_\_\_

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| **Atomic Number** | **Electronegativity** |
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* 1. In a full sentence, state the trend between groups and electronegativity.
  2. Use Bohr models for atoms in your group to represent the trend in electronegativity for groups.

1. Using your reference table S, graph the Atomic Number versus Electronegativity for Period \_\_\_\_.

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| |  |  | | --- | --- | | **Atomic Number** | **Electronegativity** | |  |  | |  |  | |  |  | |  |  | |  |  | |  |  | |  |  | |  |  | |  |  |  |  |  |  |  |  |  |
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* 1. In a full sentence, state the trend between periods and electronegativity.
  2. Use Bohr models for atoms in your period to represent the trend in electronegativity for periods.