**Vocabulary**

For each word, provide a short but specific definition from YOUR OWN BRAIN! No boring textbook definitions. Write something to help you remember the word. Explain the word as if you were explaining it to an elementary school student. Give an example if you can. Don’t use the words given in your definition!

Reaction Rate: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Enthalpy: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Entropy: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Potential Energy: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Catalyst: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Activation Energy: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Activated Complex: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Spontaneous Reaction: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

LeChatelier’s Principle: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

 Concentration/Heat: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Pressure/Volume: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

**Video 9.1 Reaction Rates**

*Key Ideas:*

* Collision theory states that a reaction is most likely to occur if reactant particles collide with the proper energy and \_\_\_\_\_\_\_\_\_\_\_\_.
* The rate of a chemical reaction depends on several factors: temperature, concentration, nature of reactants, \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_, and the presence of a catalyst.

*More Notes:*

* When the temperature of a reaction is \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ the reaction rate increases and the reaction moves faster. This is a \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ relationship.
* In order to make a reaction move faster, the concentration of the reactants can be \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_. This is a \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ relationship.
* When substances in the solid phase react, it is best to \_\_\_\_\_\_\_\_\_\_\_\_\_\_ the surface area of the reactants. This is a \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ relationship.
* When substances are gaseous, increasing the pressure of the system will \_\_\_\_\_\_\_\_ the rate of reaction because particles will be closer together and collide more often. This is a \_\_\_\_\_\_\_\_\_\_\_\_\_\_ relationship.
* Catalysts \_\_\_\_\_\_\_\_\_ the reaction rate by providing an alternate pathway for the reaction to take place.
* Ionic substances react \_\_\_\_\_\_\_\_\_\_\_\_\_\_ than covalent substances due to their ability to quickly dissolve and switch ions.

*Questions:*

1. Does every reactant collision result in a product being formed? Explain your answer.
2. How can we ensure more collisions will take place in order to increase the probability of a reaction?
	1. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
	2. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
	3. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
	4. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
	5. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
3. Refrigerated food last longer than the same food left out at room temperature. Explain.

**Video 9.2 Enthalpy Review**

*Notes*:

* Enthalpy is a fancy word for \_\_\_\_\_\_\_\_\_\_ of reaction.
* All exothermic reactions are given a \_\_\_\_\_\_\_\_\_\_\_\_\_ value on Table I.
* Exothermic reactions are \_\_\_\_\_\_\_\_ stable then endothermic reactions.
* Combustion reactions are the first 13 reaction on table I. What reactant do they have in common? \_\_\_
* The last 6 reactions on table I show substances dissolving. What reactant do they have in common? \_\_

*Questions:*

1. How can you determine if a reaction is endothermic or exothermic based on the sign of ΔH?
2. If a reaction is endothermic, is the heat a reactant or product? Explain.
3. Is the formation of aluminum oxide endothermic or exothermic?
4. Is the decomposition of nitrogen dioxide endothermic or exothermic?
5. Which is more stable: C2H6 or C2H4? Explain.
6. In the reaction where C2H2 is formed, which system has more energy, the reactants or the products?

**Fe2O3 + 3CO 🡪 2Fe + 3CO2 + 26.3KJ**

1. How much heat is produced when 3.56 moles of CO are used?
2. How many moles of Fe are produced when 356KJ of heat are made?
3. How much heat is produced when 56.8grams of Fe2O3 are used?

**Video 9.3 Potential Energy Diagram**

*Key Ideas:*

* Energy released or absorbed by a chemical reaction can be represented by a \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ \_\_\_\_\_\_\_\_\_\_\_\_\_\_ diagram.
* Energy released or absorbed during a chemical reaction (heat of reaction) is equal to the difference between the potential energy of the \_\_\_\_\_\_\_\_\_\_\_\_ and the potential energy of the \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_.
* A catalyst provides an alternate reaction pathway, which has a\_\_\_\_\_\_\_\_\_\_\_\_\_\_ activation energy than an uncatalyzed reaction.

*More Notes:*

* The \_\_\_\_\_\_\_\_\_ energy is the energy needed to start a reaction. It is labeled on the diagram with \_\_\_\_.
* The activated complex is the point with the \_\_\_\_\_\_\_\_\_ energy on the graph. It is labeled on the diagram with \_\_\_\_.
* The \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ are always at the beginning of a reaction and is labeled here with a \_\_\_\_.
* The \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ are always at the end of a reaction and is labeled here with a \_\_\_\_.
* The enthalpy of the reaction (or heat) is the \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ between the products and the reactants and is labeled here with a \_\_\_\_.
* Reactions that have products with higher energy than reactants, such as the one in the diagram, are \_\_\_\_\_\_\_thermic reactions.

1. How much energy do the reactants have?

2. How much energy do the products have?

3. Is the reaction endothermic or exothermic?

4. What is the enthalpy value?

5. How much energy does the activated complex have?

6. How much energy does this reaction need to start (activation energy)?

7. If this reaction was reversed, how much activation energy would it need?

8. If this reaction was reversed, would it be endothermic or exothermic?

9. If this reaction was reversed, what would be the value of the enthalpy of reaction?

10. Draw what a catalyst would do to this reaction’s energy on the graph.

*Questions:*



**Video 9.4 Entropy**

*Notes:*

* Entropy measures the \_\_\_\_\_\_\_\_\_\_\_\_\_\_ in a system.
* The \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ phase has the most entropy because its particles move the most.
* The more particles a system has, the \_\_\_\_\_\_\_\_\_\_ entropy it has. This is a direct relationship.
* In order for a reaction to be spontaneous it should have \_\_\_\_\_\_\_\_\_\_\_\_ entropy and be \_\_\_\_\_ thermic which has a \_\_\_\_\_\_\_\_\_ *(positive or negative)* value.

*Questions:*

1. Which substance has the highest entropy?
	1. Xe(g) b. S(s) c. Hg(s)
2. Which substance has the lowest entropy?
	1. H2O at -120˚C b. H2O at 0˚C c. H2O at 100˚C
3. When a system becomes less random the entropy
	1. Increases b. decreases c. remains the same
4. Which shows an increase in entropy?
	1. H2O(s) 🡪 H2O(l) b. H2O(g) 🡪 H2O(l) c. H2O(l) 🡪 H2O(s)
5. The following reactions shows the entropy is CaCO3(s) 🡪 CaO(s) and CO2(g)
	1. Increases b. decreases c. remains the same
6. The following reactions shows the entropy is 2AB 🡪 2A+B2
	1. Increases b. decreases c. remains the same
7. A chemical reaction is spontaneous if
	1. There is a gain of energy and entropy increases
	2. There is a gain of energy and entropy decreases
	3. There is a loss of energy and entropy increases
	4. There is a loss of energy and entropy decreases
8. If a reaction is endothermic but still spontaneous the entropy must \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

**Video 9.5 LeChatelier’s Principle**

*Key Ideas:*

* Some chemical and physical changes can reach equilibrium.
* At equilibrium the rate of the forward reaction\_\_\_\_\_\_\_\_\_\_\_\_\_\_ the rate of the reverse reaction. The measurable quantities of reactants and products remain constant at equilibrium.
* LeChatelier’s principle can be used to predict the effect of\_\_\_\_\_\_\_\_\_\_\_ (change in pressure, volume, concentration, and temperature) on a system at equilibrium.

*Notes:*

* Equilibrium states that a reversible reaction has equal \_\_\_\_\_\_\_\_\_\_\_\_\_ not necessarily equal amounts.
* Equilibrium requires the rate that the reaction forms products must be \_\_\_\_\_\_\_\_\_\_\_\_\_\_ the rate that the reaction forms reactants.
* When a reactant is added the reaction shifts away from the reactant to make more \_\_\_\_\_\_\_\_\_\_\_\_\_\_.
* When a product is added the reaction shifts away from the product to make more \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_.
* When a reactant is taken the reaction shifts \_\_\_\_\_\_\_\_\_\_\_ from the reactant to restore it.
* When a product is taken the reaction shifts \_\_\_\_\_\_\_\_\_\_\_ from the product to restore it.
* When heat is added to a reaction the reaction shifts \_\_\_\_\_\_\_\_\_\_\_\_\_\_ from the heat, much like a reactant or product. Therefore if heat is removed (or taken) the reaction shifts \_\_\_\_\_\_\_\_\_\_\_\_\_\_ heat.
* When a gaseous system increases in pressure or decreases in volume the system can hold fewer moles. Therefore the reaction shifts towards the side with \_\_\_\_\_\_\_\_\_\_ moles.

*Questions:*

1. Label where equilibrium is achieved in the system.
2. Write an equation for the system at equilibrium.
3. In the equation: **PCl5(g) + 100kJ PCl3(g) + Cl2(g)**

describe what happens to the chlorine when the following stressors are added to the system at equilibrium:

1. PCl5 is added \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
2. PCl3 is added \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
3. PCl5 is removed \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
4. PCl3 is removed \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
5. Heat is added \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
6. Heat is removed \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
7. Pressure is added \_\_\_\_\_\_\_\_\_\_\_
8. Pressure is relieved \_\_\_\_\_\_\_\_\_\_\_
9. The volume of the chamber is decreased \_\_
10. The volume of the chamber is increased \_\_
11. A catalyst is added \_\_\_\_\_\_\_\_\_\_\_