Name: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ Moles of a Colored Mole

Guiding Question: How many moles of paraffin, C25H52, are used to color your mole?

Prelab Questions:

1. What is the molar mass of the paraffin used in the crayons?
2. Another student performed this lab using chalk on a chalk board. The chalk is predominately calcium carbonate. The mass of chalk used to color the mole was 1.2grams.
	1. What is the molar mass of calcium carbonate?
	2. How many moles of calcium carbonate were used?
	3. How many molecules of calcium carbonate were used?
	4. How many atoms of calcium carbonate were used?

Analysis Questions:

1. Calculate the number of grams used to color your mole.
2. Calculate the number of moles used to color your mole.
3. Calculate the number of molecules used to color your mole.
4. Calculate the number of atoms used to color your mole.
5. Why is paraffin considered to be an organic hydrocarbon?
6. What is the name of the homologous series to which paraffin belongs?
7. Draw the structural formula of paraffin.

